

Chemistry 1A - 62 Course Outline

Fall, 2018

INSTRUCTOR: Dr. Billie Lo billielo@comcast.net

Lecture: MW 6:00 – 7:15PM S34

Laboratory: TTh 7:30 – 10:20 PM SC2202

Credit: 5 units

PREREQUISITE:

Chem. 25 with a C or better or high school chemistry with a B or better, Math C or higher.

ACCEPTABLE FOR CREDIT:

University of California, California State University and Colleges.

COURSE DESCRIPTION:

Chem. 1A is a pre-professional chemistry preparation for students planning a scientific or science related career field. A rigorous study of the fundamentals of chemistry at the first year level combines the study of atomic and molecular structure, quantum theory, thermo chemistry, gases, solutions, and qualitative analysis with the classical study of properties of atoms and molecules and their reactivity. The course includes both lecture and lab work designed to prepare students to enter as chemistry engineering, medicine, dentistry as well as biological science.

TEXTS:

Chemistry, The Molecular Nature of Matter and Change, Martin Silberberg, McGraw Hill, 8th edition, 2017.

Lab Manual

Can be found on-line at <https://www.deanza.edu/chemistry/pdf/1A/Experiments>

Click on the Experiments and download the details for each experiment.

A Simple Scientific Calculator (non-programmable) and **Safety goggles** are required.

*****Programmable calculators are not allowed for the exams or quizzes; students must use a simple scientific calculator instead.**

Academic Dishonesty: Any form of academic dishonesty will be ground for dismissal from the course.

BASIS OF EVALUATION

A. Quizzes (Approx. 5 - 10 minutes):

Quizzes will be given at the beginning or end of class meetings to those students who are present at the time of quizzes. Each quiz counts about 15 points. Pre-lectured reading materials may be covered at end of the lecture. No make-up quiz is given.

B. Hourly Exam:

Three hourly exams will be given during the semester. Make-up exam shall be given for serious and compelling reasons only. Consult your instructor PRIOR TO **EXAM TIME** by all means. There will be 10% deduction in grades for all the make-up exams.

C. Final Exam:

A comprehensive final exam will be given. Student who miss or fail the final exam will not receive a grade C or better.

D. Homework

The “Connect” **on-line homework assignments** are divided into two different parts for each chapter – the conceptual and the selected end of the chapter problems. **The advantage of doing them on-line is that you can get instant feed back when you make a mistake. You are encouraged to use the “help” or “hint” on-line to save time.** The program is set to “auto-submit” on the due day. Doing it in a timely manner would help you understand the materials better, so that you can get better grades. Feel free to open the finished assignments for review because the final performance report sum up your highest score for each chapter only. You should try to do a few problems each day. The due day is usually set right on or only a few days after the lecture on the chapter is done. On completion of 70% of the total assigned homework you will get 50 extra credit **points** toward your over-all grades. Each chapter assignment is set to open for 3 to 4 weeks and you only need to finish 70% of the total to get full credit. Therefore, usually no extension will be granted to individual student.

- E. The bookstore sells the textbook, with or without the Connect Access Code. With the /Connect access Code, it costs \$142.90, and without the Connect code costs \$107.20. You may also purchase the access Code on-line if you already have a textbook. To access “CONNECT”

For section 61 go to: <http://connect.mheducation.com/class/b-lo-61>

For section 62 go to: <http://connect.mheducation.com/class/b-lo-62>

F. Attendance:

Attendance will be enforced. **Any student who has two or more lab or lecture absences may be dropped from the course.**

G. Chemical Disposal:

As a concern for the environment, proper chemical disposal is essential. Students who do not comply with directed procedures may be dropped from the course for repeated offenses.

H. Grading:

Exams	330
Quizzes	120+
Final exam	250

Lab Grade 30%

Lab Midterm and Lab Final	140
Lab Reports	90
Lab Notebook	50
Lab Performance	20

Total 1000

880+ pts A

780+ pts B

650+pts C

500+pts D

Extra Credit: 50 Points when complete 70% of the assigned “Connect” on-line homework -

G. Worksheet schedule: Extra points

Five worksheet assignments will be given, 10 points each. Worksheets will be graded according to accuracy and neatness. Points will be deducted if late (-10% for each additional day.)

Worksheet #	Content	Chapter	Date open	Date Due
1	Nomenclature	2	10/1/18	10/8/18
2	Concentration units, Acid Base	3	10/8/18	10/22/18
3	Net Ion equations	4	10/15/18	10/22/18
4	Balance Equations	4, Expt. 9	10/17/18	10/24/18
5	Geometry (shape)	9,10	11/26/18	12/5/18

CHEMISTRY 1A TENTATIVE LECTURE AND EXAM SCHEDULE Lo**

** For section 62 Lab schedule, change M and W to T and Th; and for week 8 and 9 it will be different because of the Veterans DAY and the Thanksgiving holiday.

	CHEM 1A	LECTURE	& EXAM SCHEDULE	LABORATORY SCHEDULE
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WK	DATE	CHAPTER	CONTENT	
1	M 9/24/18	Chapter 1	Measurement, Units, Uncertainty, Precision and Accuracy, Scientific Notation	Lab Check-In
	W 9/26 /18	Chapter 1	Mathematical Treatment of Measurement Results, Atomic Theory, Atomic Structure	Measurement
2	M 10/1/18	Chapter 2	Chemical Formulas, the Periodic Table, Atomic Structure, Symbols, Nomenclature	Nomenclature
	W 10/3/18	Chapter 3	Stoichiometry of Formula and Equations	Expt 4 Hydrate(1)
3	M 10/8/18	Chapter 3	Formula Mass and the Mole Concept, Empirical Formula and Molecular Formula, Molarity and Other Units for Concentration	Expt 4 Hydrate(2) Empirical Formula of Hydrate (2)
	W 10/10/18		<i>Exam 1</i>	Precipitation(1)
4	M 10/15/18	Chapter 4	Writing and Balancing Chemical Equat'ns, Classifying Chemical Reactions	Precipitation(2)
	W 10/17/18	Chapter 4	Reaction Stoichiometry, Yields, Quantitative	Precipitation(3)
5	M 10/22/18	Chapter 6	Thermochem: Hess's Law, Calorimetry, Enthalpy	Type of reactions(1)
	W 10/24/18	Chapter 6	Thermochemistry: Calorimetry, Enthalpy	Type of reactions(2)
6	M 10/29/18	Chapter 6	<i>Exam 2</i>	Lab Midterm / Conductivity (1)
	W 10/31/18	Chapter 7	Radiation- Energy, Electromagnetic Waves, the Bohr Model	Conductivity(2)
7	M 11/5/18	Chapter 7	Quantum Theory, Quant # & sublevel-orbitals	Expt 5 Acid-Base Titration (1)
	W 11/7/18	Chapter 8	Electron Configuration & Chem. Periodicity (Trends in Ionization Energies, Electronegativities)	Expt 5 Acid-Base Titration (1)
8	M 11/12/18		VETARERANS DAY HOLIDAY	
	W 11/14/18		<i>Exam 3</i>	Expt 6 Calorimetric Msurm'ts 1)
	11/16/18		<i>Last Day to drop the class with a "W"</i>	
9	M 11/19/18	Chapter 8	Electron Configuration & Chem. Periodicity (Trends in Ionization Energies, Electronegativities)	Expt 6 Calorimetric Measurements (1)
	W 11/21/18	Chapter 9	Models of Chemical Bonding	Expt 9 Redox Titration (1)
	Nov. 22 –	Nov. 25	Thanksgiving Holiday	
10	M 11/26/18	Chapter 10	Molecular Structure, VSEPR Theory, Shape and Polarity,	Expt 9 Redox Titration (2)
	W 11/28/18	Chapter 11	Molecular Structure, VSEPR Theory, Shape and Polarity,	Expt 10 Line Spectra of H-atom Expt 11 Molecular Model (1)
11	M 12/3/18	Chapter 11	Valence Bonds Theory, Hybrid Atomic Orbitals, Multiple Bonds, Molecular Orbital Theory	Expt 11 Molecular Model (2)
	W 12/5/18	Review		Lab final /Check - Out
12	M 12/10/18	Final Exam		

J. CHEMISTRY 1A LABORATORY

1. **SAFETY GLASSES OR GOGGLES** must be worn **AT ALL TIMES** while you are in the laboratory.
2. Each student is required to have a **lab notebook** to outline the lab procedures, record experiment data, and calculations. It will be evaluated as part of the lab grade.
3. You are expected to arrive in the laboratory on time. Tardiness of 15 minutes or more will not be permitted. Preview the lab materials before coming to lab is required
4. Students must clean and return all items borrowed to the stock room **no later than 10:05 PM** each day of the experiment.
5. Student must check out with the instructor at the end of each lab to have their notebook stamped and sign a roll sheet.
6. Each laboratory experiment must be completed within the specified time. When that period is over, no credit will be given for the lab, but **all labs must be completed to receive a grade in the course**. All lab work not conducted will be graded as zero.
7. **Chemical Disposal:**
Proper chemical disposal is essential. Students who do not comply with directed procedures may be dropped from the course for repeated offenses.
8. Please note that you are required to **officially** check out of your lab locker whether you remain in the course or drop the course. Failure to check out of lab on time will result in a late fee and may also result in your grades being held and a block placed on your future registration.
9. **If you drop within the first two weeks of class and fail to check out of lab, your locker may be reassigned to another student by the instructor, and you will be held responsible for any missing or broken lab locker equipment. After the first two weeks of class you must checkout by the assigned checkout date for your laboratory section.**

K. FORMAT OF THE LABNOTEBOOK (a permanently bound notebook): *Everything should be written in ink in the notebook.*

1. Number and Title of the experiment
2. Purpose/theory of the experiment (brief)
3. Formula for the calculation.
4. Procedure in detail for the experiment. A photocopy of the lab manual is not allowed. Check with the lab instructor which section will be performed next to minimize preparation time and effort.

The above should be fully prepared prior to attending the lab lecture and it should be stamped before lab lecture.

5. Data (laboratory work) must be entered **immediately** and **directly** into the lab notebook **in ink**.
6. Calculations: The laboratory midterm and final are **"open-notebook"**. A well-prepared notebook would be helpful during these exams.

L. FORMAT OF THE LAB REPORT – One lab report per each experiment. No point will be given Unless the student actually performed the experiment.

1. Number and Title of the experiment.
2. Theory (more detail) and formula for the calculation
3. Procedure for the experiment (brief).
4. Data and calculation. Show at least one set-up for each different type of calculations.
5. Results (including all graphs) and discussion in doubt.
6. Pre-lab and post-lab questions must be answered and attached to the lab report.

Report is due on day 2 of the next experiment. Penalty for late reports: 1-2 day late less 10%, 2-7 day late less 40% More than 1 week late, less 60%.

Student Learning Outcome(s):

*Identify and explain trends in the periodic table.

*Construct balanced reaction equations and illustrate principles of stoichiometry.

*Apply the first law of thermodynamics to chemical reactions.